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Gases in a Nonflexible Container

Experiment A (Adding more gas) 25 Gas

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Variables-S Just remember the general gas law  $\mathrm{PV=nRT}$  or rearranging terms we see that  $\mathrm{\frac{[PV]}{[nT]}=R}$  where R is a constant.

## **Gas Variables Answers**

Model 1 - Gases in a Nonflexible  
Container Experiment A (Adding more

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gas) A1 A2 A3 Volume = 1 unit Volume  
= 1 unit Volume = 1 unit External  
pressure = 1 atm External pressure = 1  
atm External pressure = 1 atm Internal  
pressure = 1 atm Internal pressure = 2  
atm Internal pressure = 3 atm  
Temperature = 200 K Temperature =  
200 K Temperature = 200 K Experiment  
B (Heating the gas) B1 B2 B3 Volume =

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1 unit Volume = 1 unit Volume = 1 unit  
External pressure = 1 atm External  
pressure = 1 atm ...

## **Model 1 Gases in a Nonflexible Container Experiment A ...**

(V), temperature (T), and moles (n) of  
gas. These four variables can be related  
mathematically so that predictions about



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gas behavior can be made. Model 1 -  
Gases in a Nonflexible Container  
Experiment A (Adding more gas) A1  
A3A2 Volume = 1 unit Volume = 1 unit  
Volume = 1 unit

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key What happens when gas in a  
container is heated - The Q&A â€¦ ...

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All gases must be enclosed in a  
container that, if there are openings, can

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be sealed with no leaks. The three-dimensional space enclosed by the container walls is called volume. When the generalized variable of volume is discussed, the symbol  $V$  is used. Volume in chemistry is usually measured in liters (symbol = L) or milliliters (symbol = mL).

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## **ChemTeam: Four Gas Law Variables**

More gas is then added to the container until it reaches a final volume of 13.5 L. Assuming the pressure and temperature of the gas remain constant, calculate the number of moles of gas added to the container. Solution: 1) Let's start by rearranging the Ideal Gas Law (which you'll see a bit later or you can go

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review it right now):  $PV = nRT$

## **ChemTeam: Gas Law - Avogadro's Law**

add a lot of gas? In this activity, you will explore four variables that quantify gases—pressure (P), volume (V), temperature (T), and moles (n) of gas. These four variables can be related

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mathematically so that predictions about gas behavior can be made. Model 1 - Gases in a Nonflexible Container Experiment A (Adding more gas) A1 A3A2

## **POGIL Chemistry Activities - Flinn**

Gas molecules keep their distance from each other and are in constant motion.

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They continue to move in one direction until they come into contact with an object. Gas expands when placed in a closed container. The molecules continue to move about, filling the container. They strike the sides of the container, and each ...

## **What Three Factors Affect the**

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## **Pressure of the Gas in a ...**

In a non flexible container as the volume increases the more space the gas has to move around and there is less collisions resulting in a decrease in pressure.

Oppositely, as the volume decreases the molecules have less space to move and the molecules collide more often resulting in more pressure.



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## **Chemistry Test Review Flashcards | Quizlet**

A nonflexible container is needed if the gas sample is going to have an internal pressure that is different from the external pressure. If a flexible container is used, the internal pressure and external pressure will always be the

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same because they are both pushing on the sides of the container equally.

## **25 Gas Variables-S - Weebly**

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## **Gas Variables How Are The Variables That Describe A Gas ...**

Answer to what two materials ensure that containers with gas are nonflexible? chemistry...

## **Solved: What Two Materials Ensure**

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## **That Containers With Gas ...**

Of the variables that were controlled in both Experiment A and Experiment B in Model 1, one requires a nonflexible container. Name this variable, and explain why a nonflexible container is necessary. In your answer, consider the external and internal pressure data given in Model 1.

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## **Eraj Khan - Gas Variables HW.pdf - Gas Variables How are ...**

This is an example where the pressure inside the closed container is higher than atmospheric pressure, but it can be lower as well. What happens to the gas in a container depends on how flexible the container is. If I take a glass bottle

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and pump half the air out, the gas inside will stay at half atmospheric pressure.

### **Why does a gas inside a closed container experience ...**

See the answer 1. The gas inside a flexible container (e.g., a balloon) is allowed to expand against a constant external pressure of 2.1 bar until the

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volume increases from 2.7 L to 8.6 L.  
Calculate the amount of work done on  
the system.

## **Solved: 1. The Gas Inside A Flexible Container (e.g., A Ba ...**

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Volume and pressure in gases - the gas laws Boyle's law. Decreasing the volume of a gas increases the pressure of the gas. An example of this is when a gas is



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trapped in a cylinder by a piston.

## **Volume and pressure in gases - the gas laws - Temperature ...**

Yes, a gas will always have the volume of its container. The volume will not change if you change the number of molecules, what will change is the pressure. There are some great videos

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where they heat water inside a 55 gallon drum. The water turns...

### **Is the volume of a gas equal to the volume of its ... - Quora**

nonflexible container is necessary. In your answer, consider the external and internal pressure data given in Model 1. Name the two factors related to

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molecular movement that influence the pressure of a gas. For each experiment in Model 1, determine the relationship between the independent and dependent variables, and write an algebraic expression for the relationship using

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